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~~Ch 14 Acids and~~

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Bases and Salts

CH 14 Acids and

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~~Newton's Laws of~~

~~Motion~~

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Naming Acids

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Acids + Bases

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Heck is an Acid

or Base? -

Organic

Chemistry pH and

pOH: Crash

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#30 Acids and

Bases and Salts

— Introduction —

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~~Chemistry |~~

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Kinetics) - Part

1 Defects of

Vision and Their

Correction Acids

and Bases

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Chapter 26

Fluid,

Electrolyte,

Acid-Base

Balance

Chapter 16 (Acid-
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- Part 1 K_a K_b

K_w pH pOH pKa

pKb H^+ OH^-

Calculations -

Acids \u0026amp;

Bases, Buffer

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Solutions And

Chemistry Review

Acids and Bases

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and Bases

Lecture#1 Part#1

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bases part 3

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and Bases Video

2 of 3 ~~Chapter~~

~~14 Acids And~~

~~Bases~~

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Chapter 14

Acids and Bases

1. Chapter 14:
Acids and Bases

1. Chapter 14.

Acids and Bases.

1. Which of the following is the equilibrium constant

expression for the dissociation of the weak acid HOCl? a) $K =$ b) K

$= [H^+][OCl^-]$ c) K

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= d) K = e) none
of these. 2. In
which of the
following

reactions does
the H_2PO_4^- ion
act as an acid?

a) $\text{H}_3\text{PO}_4 + \text{H}_2\text{O}$
 $\text{H}_3\text{O}^+ + \text{H}_2\text{PO}_4^-$

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~~Acids and Bases~~

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Acids and Bases;

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Kids and

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Examples; 14.3:

Bases:

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Examples; 14.4:

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14.5: Reactions
of Acids and
Bases; 14.6:
Acid-Base

Titration: A Way
to Quantify the
Amount of Acid
or Base in a
Solution; 14.7:
Strong and Weak
Acids and Bases

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and Bases

Arrhenius

Concept: Acids

produce H^+ in

solution, bases

produce OH^- ion.

Br nsted-Lowry:

Acids are H^+

donors, bases

are proton

acceptors.

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and Bases And

Introduction to
Acids and Bases

• Acids : •

Carry • Have • D

issolve • Turn 2

H got tons (

hydro ni um ion

) a sour taste

many metals blue

litmus paper red

Introduction to

Acids and Bases

• Acids : •

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Examples: H_2

SO_4^{2-} $\text{HC}_2\text{H}_3\text{O}_2$

HNO_3

HCl H_3PO_4

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~~Introduction...~~

General

Properties of
Acids.

Definition. 1)

sour taste,

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corrosive. 2) characteristically changes acid-base indicator.

3) react with active metals to form hydrogen gas. 4) react with bases to form salt and water. 5)

conduct electric current. Term.

General

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Properties of
Acids And
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October 10th,
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some maintenance
on Quizlet to
keep things
running
smoothly.

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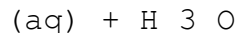
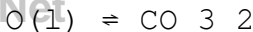
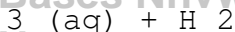
Acids are proton
(H^+) donors, and
bases are proton

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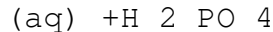
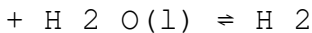
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as an acid: HCO_3^-

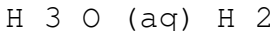
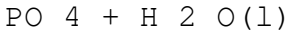


+ (aq) HCO_3^- as a

base: $\text{HCO}_3^- (\text{aq})$



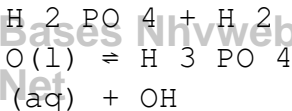
as an acid: H_2PO_4^-



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PO₄ as a base:



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Acids and Bases.
Know the
definition of
Arrhenius,

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Bronsted-Lowry,
and Lewis acid
and base.

Autoionization
of Water. Since
we will be
dealing with
aqueous acid and
base solution,
first we must
examine the
behavior of
water. 1) In
pure water, to

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which no electrolytes have been added, there is a small electrical conductivity, indicating that there are ions present.

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Base Pairs,

Arrhenius,

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Acids and Bases.

binary acid.

oxyacid.

Arrhenius acid.

Arrhenius base.

an acid that

contains only

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two different
elements:

hydrogen an... an
acid that is a
compound of
hydrogen,
oxygen, and a
third el... a
chemical
compound that
increases the
concentration of
hydrog...

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And Bases

Worksheet

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14.1 The Nature

of Acids and

Bases. A.

Page 30/85

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Arrhenius Model

1. Acids produce hydrogen ions in aqueous

solutions 2.

Bases produce hydroxide ions in aqueous

solutions B.

Bronsted-Lowry

Model 1. Acids

are proton

donors 2. Bases

are proton

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acceptors 3.

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1 Chapter 14:

Acids, Bases and

Salts Sections

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nature of acids

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and bases And

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Bases Nitroweb
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The pH scale

14.4)

Calculating the
pH of strong
acid solutions

14.5)

Calculating the
pH of weak acid
solutions

14.6) Bases

14.7)

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19 - Day A

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bases , and pH

scale 2) Read

14.1- 14.4

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467 SECTION 1 O

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five general properties of aqueous acids and bases. Name common binary acids and oxyacids, given their chemical formulas. List five

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